

Answers to 1/3 sheet writing/naming chemical formulas worksheet:

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|----|--|---|--|
| 2) | a) K_2S
d) Fe_2O_3
g) Na_2CO_3 | b) $Ca(OH)_2$
e) $CuCl_2$
h) Ti_2O_3 | c) $(NH_4)_2SO_4$
f) $AgNO_3$ |
| 3) | a) zinc sulfide
c) hydrogen sulfate
e) sodium iodide
g) silver sulfide
i) chromium sulfate | b) potassium nitrate
d) ammonium hydroxide
f) sodium oxide
h) magnesium sulfate
j) sulfur dioxide | |
| 4) | a) $NiBr_3$
d) $Al_2(CO_3)_3$
g) P_2O_3
j) calcium carbonate | b) Ag_2O
e) NO_3
h) silicon dioxide
k) magnesium oxide | c) O_2
f) CCl_4
i) hydrogen chloride
l) magnesium carbonate |

Answers to the ½ sheet review sheet!!!!

- 1)
 - a) only has two atoms or polyatomic ions in it; ends in -ide
 - b) bond where atoms transfer electrons; involves a metal and a nonmetal
 - c) bond where atoms share electrons
 - d) a group of atoms that behave like a single atomic ion
 - e) elements in group 18; they have full outer shells or eight electrons in the outer shell; they do not typically react since they satisfy the octet rule
 - f) the charge on the atom after it gains or loses electrons
- 2) a metal and a nonmetal, or possibly a polyatomic ion instead of one of them
- 3) nonmetals only
- 4) Only with covalent compounds... when there are only nonmetals involved. The prefixes are:

Mono – 1	hexa – 6
Di – 2	hepta – 7
Tri – 3	octa – 8
Tetra – 4	nona – 9
Penta – 5	deca – 10
- 5) This is the charge on the atom (the ion) after it gains or loses electrons.
- 6) Count the number of valence electrons. If it has 3 or less, it will lose those electrons, and if it has 5, 6, or 7, it will gain electrons until it has 8. Count the number gained (that