

## **Naming Ionic Compounds Practice Worksheet**

*Name the following ionic compounds:*

- 1)  $\text{NH}_4\text{Cl}$  \_\_\_\_\_
- 2)  $\text{Fe}(\text{NO}_3)_3$  \_\_\_\_\_
- 3)  $\text{TiBr}_3$  \_\_\_\_\_
- 4)  $\text{Cu}_3\text{P}$  \_\_\_\_\_
- 5)  $\text{SnSe}_2$  \_\_\_\_\_
- 6)  $\text{GaAs}$  \_\_\_\_\_
- 7)  $\text{Pb}(\text{SO}_4)_2$  \_\_\_\_\_
- 8)  $\text{Be}(\text{HCO}_3)_2$  \_\_\_\_\_
- 9)  $\text{Mn}_2(\text{SO}_3)_3$  \_\_\_\_\_
- 10)  $\text{Al}(\text{CN})_3$  \_\_\_\_\_

*Write the formulas for the following compounds:*

- 11) chromium (VI) phosphate \_\_\_\_\_
- 12) vanadium (IV) carbonate \_\_\_\_\_
- 13) tin (II) nitrite \_\_\_\_\_
- 14) cobalt (III) oxide \_\_\_\_\_
- 15) titanium (II) acetate \_\_\_\_\_
- 16) vanadium (V) sulfide \_\_\_\_\_
- 17) chromium (III) hydroxide \_\_\_\_\_
- 18) lithium iodide \_\_\_\_\_
- 19) lead (II) nitride \_\_\_\_\_
- 20) silver bromide \_\_\_\_\_