

Nomenclature Worksheet #4 – Polyatomic Compounds

WRITE FORMULAS FOR EACH OF:

- | | |
|---------------------------------|---|
| 1. lithium nitrate _____ | 17. aluminum sulfate _____ |
| 2. sodium chlorate _____ | 18. aluminum phosphate _____ |
| 3. calcium nitrate _____ | 19. copper (I) sulfate _____ |
| 4. sodium phosphate _____ | 20. copper (I) phosphate _____ |
| 5. aluminum nitrate _____ | 21. zinc sulfate _____ |
| 6. potassium sulfate _____ | 22. cobalt (III) chloride _____ |
| 7. ammonium nitrate _____ | 23. calcium hydroxide _____ |
| 8. ammonium oxide _____ | 24. sodium hydroxide _____ |
| 9. potassium chlorate _____ | 25. aluminum hydroxide _____ |
| 10. barium carbonate _____ | 26. ammonium hydroxide _____ |
| 11. iron (II) sulfate _____ | 27. sodium bicarbonate _____ |
| 12. iron (II) nitrate _____ | 28. ammonium bicarbonate _____ |
| 13. iron (II) phosphate _____ | 29. copper (II) bicarbonate _____ |
| 14. titanium (IV) sulfate _____ | 30. aluminum bicarbonate _____ |
| 15. nickel (III) nitrate _____ | 31. iron (III) hydrogen carbonate _____ |
| 16. silver nitrate _____ | 32. sulfate radical _____ |

NAME EACH OF THE FOLLOWING:

33. CaSO_4 _____
34. LiOH _____
35. Ba(OH)_2 _____
36. KNO_3 _____
37. Pb(OH)_2 _____
38. $\text{Ba}_3(\text{PO}_4)_2$ _____
39. $(\text{NH}_4)_2\text{SO}_4$ _____
40. $\text{Mg(NO}_3)_2$ _____
41. $\text{Ni}_2(\text{SO}_4)_3$ _____
42. $\text{Pb(ClO}_3)_2$ _____
43. Na_2SO_4 _____
44. SrCO_3 _____