

Molarity Calculations – Answer Key

Calculate the molarities of the following solutions:

- 1) 2.3 moles of sodium chloride in 0.45 liters of solution. 5.11 M
- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution. 0.98 M
- 3) 0.09 moles of sodium sulfate in 12 mL of solution. 7.5 M
- 4) 0.75 moles of lithium fluoride in 65 mL of solution. 11.5 M
- 5) 0.8 moles of magnesium acetate in 5 liters of solution. 0.16 M
- 6) 120 grams of calcium nitrite in 240 mL of solution. 3.79 M
- 7) 98 grams of sodium hydroxide in 2.2 liters of solution. 1.11 M
- 8) 1.2 grams of hydrochloric acid in 25 mL of solution. 1.35 M
- 9) 45 grams of ammonia in 0.75 L of solution. 3.53 M

Explain how you would make the following solutions.

- 10) 2 L of 6 M HCl Dissolve 426 g HCl, dilute to 2 L
- 11) 1.5 L of 2 M NaOH Dissolve 120 g NaOH, dilute to 1.5 L
- 12) 0.75 L of 0.25 M Na_2SO_4 Dissolve 26.64 g Na_2SO_4 , dilute to 0.75 L
- 13) 45 mL of 0.12 M sodium carbonate Dissolve 0.57 g Na_2CO_3 , dilute to 45 mL
- 14) 250 mL of 0.75 M lithium nitrite Dissolve 9.92 g LiNO_2 , dilute to 250 mL
- 15) 56 mL of 1.1 M iron (II) phosphate Dissolve 22.02 g $\text{Fe}_3(\text{PO}_4)_2$, dilute to 56 mL
- 16) 6.7 L of 4.5 M ammonium nitrate Dissolve 2412 g NH_4NO_3 , dilute to 6.7 L
- 17) 4.5 mL of 0.05 M magnesium sulfate Dissolve 0.02709 g MgSO_4 , dilute to 4.5 mL
- 18) 90 mL of 1.2 M BF_3 Dissolve 7.32 g BF_3 , dilute to 90 mL