

CHAPTER 6 REVIEW WORKSHEET

Name: _____ Block: _____ Date: _____

1. What type of chemical reaction involves two smaller molecules reacting to produce one large molecule?

- A. Decomposition
- B. Synthesis**
- C. Single Replacement
- D. Combustion

2. Given the incomplete equation of a chemical reaction: $C_9H_6O_4 + O_2 \rightarrow$
Which of the following are products formed from this reaction?

- I. H_2
- II. CO_2
- III. H_2O

- A. I and II only
- B. I and III only
- C. II and III only**
- D. I, II and III

3. Which of the following represents a single replacement reaction?

- I. $Sn + 2AgNO_3 \rightarrow Sn(NO_3)_2 + 2Ag$
 - II. Gold (II) cyanide + zinc \rightarrow gold + zinc cyanide
 - III. Magnesium iodide reacts with bromine gas to produce magnesium bromide and iodine
- A. I and II only
 - B. I and III only
 - C. II and III only
 - D. I, II and III**

4. Sodium nitrate is produced as a result of mixing a solution of cadmium(II) nitrate with a solution of sodium sulphide. What is the other compound formed from this reaction?

- A. CdS
- B. $CdSO_4$
- C. NaS_2
- D. $CdNO_4$**

5. What type of reaction is represented by the following word equation:

Potassium chlorate \rightarrow Potassium chloride + oxygen

- A. Synthesis
- B. Decomposition**
- C. Single replacement
- D. Double replacement

6. Predict the products of the following chemical reactions, classify the reaction and balance it.

Reactants		Products	Reaction Type
$2NaI + F_2$	\rightarrow	$2NaF + I_2$	Single Replacement
$Cl_2 + 2KI$	\rightarrow	$2KCl + I_2$	Single Replacement
$3AgNO_3 + Na_3PO_4$	\rightarrow	$Ag_3PO_4 + 3NaNO_3$	Double Replacement
$2CH_3OH + O_2$	\rightarrow	$2CO_2 + H_2O$	Combustion
$MgCl_2$	\rightarrow	$Mg + Cl_2$	Decomposition
$3Sr(OH)_2 + 2H_3PO_4$	\rightarrow	$Sr_3(PO_4)_2 + 6H_2O$	Neutralization

