

binary ionic compounds (Homework)

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1.
A simple ion with a +1 charge (for example, Na^+) results when an atom _____ electrons.
2.
Positive ions are called _____.
whereas negative ions are called _____.
3.
Simple negative ions formed from single atoms are given names that end with the letters _____.
4.
How many electrons are contained in each of the following ions?

(a) K^+	(b) Mn^{2+}	(c) Co^{3+}	(d) Co^{2+}
(e) Cr^{3+}	(f) I^-	(g) Fe^{3+}	(h) P^{3-}
5.
For the following pairs of ions, use the concept that a chemical compound must have a net charge of zero to predict the formula of the simplest compound that the ions are most likely to form. (Type your answer using the format CO_2 for CO_2 .)

(a) Na^+ and C^{4-}	(b) Sn^{4+} and N^{3-}	(c) Fe^{3+} and P^{3-}	(d) Sn^{4+} and C^{4-}
(e) Mg^{2+} and N^{3-}	(f) Mg^{2+} and C^{4-}	(g) Fe^{3+} and C^{4-}	(h) Fe^{3+} and S^{2-}
6.
Write the correct formula for the ionic compound composed of potassium and iodide. (Type your answer using the format CH_4 for CH_4 .)
- 7.