

## Percent Composition and Molecular Formula Worksheet

1. What's the empirical formula of a molecule consisting 65.0% carbon, 5.7% hydrogen, and 29.3% oxygen?
2. If the molar mass of the compound is 108.1 g/mol, what's the molecular formula?
3. What's the empirical formula of a molecule consisting 65.7% carbon, 6.7% hydrogen, and 27.6% oxygen?
4. If the molar mass of the compound is 108.1 g/mol, what's the molecular formula?

Write the molecular formula of the following compounds:

5. A compound with an empirical formula of  $C_2H_4O$  and a molar mass of 88 g/mol.
6. A compound with an empirical formula of  $C_2H_4O$  and a molar mass of 176 g/mol.
7. A compound with an empirical formula of  $C_2H_4O$  and a molar mass of 220.7 g/mol.
8. A compound with an empirical formula of  $C_2H_4O$  and a molar mass of 88 g/mol.

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.
10. The molar mass of a general  $CH_2O$  was determined by experiment to be 60.0 g/mol. What is the molecular formula?
11. Acetic acid reacting with iron(III) oxide yields iron, carbon dioxide, and water. Combustion of such compounds yields  $CO_2$ ,  $H_2O$ , and  $N_2$  as products. If the combustion of 9.71 g of acetic acid yields 10.2 g  $CO_2$  and 1.46 g  $H_2O$ , what is its empirical formula?
12. The molar mass of acetic acid is 60 g/mol. What is its molecular formula?
13. Calculate the mass percent of carbon, hydrogen, and oxygen in acetic acid,  $C_2H_4O_2$ .
14. 6.363 g sample of a compound made from glucose and vitamin C decomposed. Analysis of the products identified 11.51 g of glucose and some water produced. What is the empirical formula of the compound?
15. When 2.000 g of an oxide of nitrogen ( $NO_x$ ) is decomposed into the elements by heating, 1.451 g of nitrogen are produced. What is the empirical formula?
16. The compound contains the following percent composition. What is the empirical formula?  
 $C = 58.81\% \text{ H} = 5.21\% \text{ N} = 35.98\% \text{ O} = 1.00\% \text{ S} = 4.99\%$
17. A compound of general molar mass has an approximate molar mass of 100 g/mol. If the percent composition is as follows, what is the empirical and molecular formula of acetic?  
 $C = 34.8\% \text{ H} = 4.64\% \text{ N} = 58.56\% \text{ O} = 0.99\%$