

Percent Composition and Molecular Formula Worksheet

1. What's the empirical formula of a molecule consisting 55.7% carbon, 5.7% hydrogen, and 38.6% oxygen?
2. If the molar mass of the compound is 118 g/mol, what's the molecular formula?
3. What's the empirical formula of a molecule consisting 55.7% carbon, 5.7% hydrogen, and 38.6% oxygen?
4. If the molar mass of the compound is 118 g/mol, what's the molecular formula?

Write the molecular formula of the following compounds:

5. A compound with an empirical formula of C_2H_4O and a molar mass of 88 g/mol.
6. A compound with an empirical formula of C_2H_4O and a molar mass of 176 g/mol.
7. A compound with an empirical formula of C_2H_4O and a molar mass of 220 g/mol.
8. A compound with an empirical formula of C_2H_4O and a molar mass of 88 g/mol.

Answer the following questions:

9. The percentage composition of a certain acid is found to be 39.7% C, 4.7% H, and 55.6% O. Determine the empirical formula of the acid.
10. The molar mass of a certain acid was determined by experiment to be 118 g/mol. What is the molecular formula?
11. A certain starting material for antibiotic synthesis consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 0.71 g of a certain compound yields 0.9 g CO_2 and 0.45 g H_2O , what is its empirical formula?
12. The molar mass of a certain acid is 118 g/mol. What is its molecular formula?
13. Calculate the mass percent of carbon, hydrogen, and oxygen in acetic acid, $C_2H_4O_2$.
14. 0.1612 g sample of a compound made from glucose and ribitol is decomposed. Analysis of the products shows that 0.110 g of glucose was produced. What is the empirical formula of the compound?
15. When 1.0000 g of an oxide of nitrogen (NO_x) is decomposed into the elements by heating, 0.400 g of nitrogen are produced. What is the empirical formula?
16. The compound contains the following percent composition. What is the empirical formula?
C = 39.7% H = 4.7% N = 55.6% O = 11.0% N = 11.0% N = 11.0% N = 11.0%
17. A compound of a certain acid has an approximate molar mass of 118 g/mol. If the percent composition is as follows, what is the empirical and molecular formula of the acid?
C = 39.7% H = 4.7% N = 55.6% O = 11.0% N = 11.0% N = 11.0% N = 11.0%