

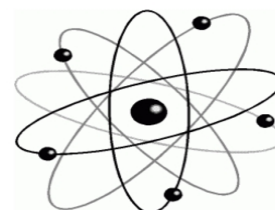
Chapter 4 Review Worksheet

Name: _____ Block: _____ Date: _____

- Which of the following is the smallest particle of an element that can exist by itself?
 - Ion
 - Atom**
 - Molecule
 - Compound

- Which of the following correctly matches the subatomic particle with its charge and location in an atom?

	Subatomic Particle	Location	Charge
A	Proton	Nucleus	Neutral
B	Neutron	Nucleus	Positive
C	Electron	Shell	Positive
D	Electron	Shell	Negative



- Which of the following are responsible for bonding?
 - Nuclei
 - Protons
 - Neutrons
 - Electrons**
- How do you calculate the number of neutrons in an atom's nucleus?
 - Atomic number
 - Mass number – atomic number**
 - Mass number + atomic number
 - Number of electrons + Number of protons

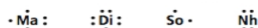
- Which of the following describes a cation?

I.	Examples include Ca ²⁺ and Al ³⁺
II.	A metal atom that has lost electrons
III.	Has an equal number of protons and electrons

- I and II only**
- I and III only
- II and III only
- I, II and III

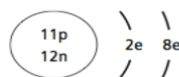
- Which of the hypothetical elements shown below represents a metal?

- Ma
- Di
- So**
- Nh



- Which of the following does the Bohr model represent?

- A neon atom
- A sodium atom
- A sodium ion**
- A fluorine atom



- Draw Bohr diagrams for the following elements:

a) Nitrogen atom



b) Potassium ion



- Draw Lewis diagrams to show the following chemical bonds:

a) CaCl₂ :



b) CO₂ :

