

Electron Configuration Worksheet

- 1) **What is meant by electron configuration?**
- 2) **Why is there a distinct electron arrangement for each atom?**
- 3) **Electrons within atoms seek out the highest energy levels (True or False).**
- 4) **What does the Aufbau principal state?**
- 5) **In which principal energy level do the different sublevels begin to overlap?**
- 6) (a) **One of the rules requires placing as many single electrons in ____.**
(b) **In this way, electron-electron repulsion is ____.**
- 7) **What is Hund's rule?**
- 8) **What is Pauli's exclusion principle?**
- 9) **How is an unoccupied orbital represented in orbital notation?**
- 10) **Show the orbital notation for the element C.**
- 11) **How are electrons represented in an electron configuration?**
- 12) **Write the electron configuration for the element C.**
- 13) **Electron dot notation shows only ____.**
- 14) **How does one recognize the highest occupied energy level?**
- 15) **What are inner shell electrons?**
- 16) **Atoms which have the s and p sublevels of their highest main energy level filled with eight electrons are said to have a(n) ____.**
- 17) **What is a noble gas configuration?**