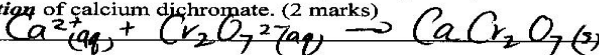
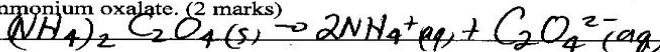


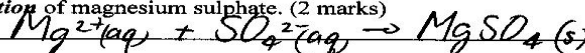
2. Write a net ionic equation (balanced, all charges, subscripts, arrow(s) correct, correct order etc.) for the **crystallization** of calcium dichromate. (2 marks)



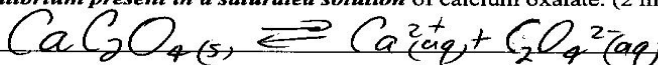
8. Write a net ionic equation (balanced, all charges, subscripts, arrow(s) correct, correct order etc.) for **dissolving** ammonium oxalate. (2 marks)



9. Write a net ionic equation (balanced, all charges, subscripts, arrow(s) correct, correct order etc.) for the **precipitation** of magnesium sulphate. (2 marks)



10. Write a net ionic equation (balanced, all charges, subscripts, arrow(s) correct, correct order etc.) for the **equilibrium present in a saturated solution** of calcium oxalate. (2 marks)



11. Calcium fluoride has a solubility of 6.87 grams/L at a certain temperature. Express this solubility in moles per Litre. (2 marks) (Show all work. Include units in your answer. Use correct # of S.D.'s)

$$\frac{6.87 \text{ g}}{\text{L}} \times \frac{1 \text{ mol}}{78.1 \text{ g}} = 0.0880 \text{ mol/L} \quad (8.80 \times 10^{-2} \text{ mol/L})$$

Answer 0.0880 mol/L

12. The molar solubility of Ag_2CO_3 at a certain temperature is $8.3 \times 10^{-5} \text{ M}$. Express this solubility in grams per Litre. (2 marks) (Show all work. Include units in your answer. Use correct # of S.D.'s)

$$8.3 \times 10^{-5} \frac{\text{mol}}{\text{L}} \times \frac{275.8 \text{ g}}{1 \text{ mol}} = 0.023 \text{ g/L} \quad (2.3 \times 10^{-2} \text{ g/L})$$

Answer 0.023 g/L

13. 0.0021 grams of MgCO_3 will dissolve in 1.0 L of water at a certain temperature. Express this solubility in grams/100 mL of water. (2 marks) (Show all work. Include units in your answer. Use correct # of S.D.'s)

$$\frac{14}{14}$$

Answer 0.00021 g/100 mL H_2O

$$(2.1 \times 10^{-4} \text{ g/100 mL } \text{H}_2\text{O})$$