CH 8 Worksheet Name: KEY

Chem 1025
Summer 2009

1. Balance the following chemical equations:

a. $2\ H_2$ + O_2 → $2\ H_2O$

b. $Pb(NO_3)_2$ + $2\ KI$ → PbI_2 + $2\ KNO_3$

c. $3\ Ca(C_2H_3O_2)_2$ + $2\ H_3PO_4$ → $Ca_3(PO_4)_2$ + $6\ HC_2H_3O_2$

d. H_2SO_4 + $2\ NaOH$ → Na_2SO_4 + $2\ H_2O$

e. $2\ HNO_3$ + $Ba(OH)_2$ → $Ba(NO_3)_2$ + $2\ H_2O$

f. $2\ Cu$ + O_2 → $2\ CuO$

g. Na_2CrO_4 + $BaCl_2$ → $BaCrO_4$ + $2\ NaCl$

h. $2\ Al$ + $3\ ZnSO_4$ → $Al_2(SO_4)_3$ + $3\ Zn$

i. $4\ Cr$ + $3\ O_2$ → $2\ Cr_2O_3$

j. $4\ P$ + $5\ O_2$ → $2\ P_2O_5$

k. $2\ N_2$ + $3\ O_2$ → $2\ N_2O_3$

l. $2\ Cl_2$ + O_2 → $2\ Cl_2O$

m. $2\ Na$ + Cl_2 → $2\ NaCl$

n. $2\ Al$ + $3\ S$ → Al_2S_3

o. $2\ Cr$ + $3\ Cl_2$ → $2\ CrCl_3$

p. $2\ Na$ + $2\ H_2O$ → $2\ NaOH$ + H_2

q. Cu + $2\ AgNO_3$ → $2\ Ag$ + $Cu(NO_3)_2$