

Solution Stoichiometry Worksheet #2

1. Consider the reaction:  $\text{Pb}(\text{NO}_3)_2(aq) + 2\text{KI}(aq) \rightarrow \text{PbI}_2(s) + 2\text{KNO}_3(aq)$ . Suppose you mix 10.0 mL of 0.20 M  $\text{Pb}(\text{NO}_3)_2$  and 20.0 mL of 0.20 M KI together. What is the concentration of  $\text{KNO}_3$  produced? What are the concentrations of the reactants that remain after the reaction goes to completion?

30.0 mL = total volume

$$0.0100\text{L} \cdot \frac{0.20\text{ mol Pb}(\text{NO}_3)_2}{1\text{L Pb}(\text{NO}_3)_2} = \frac{2\text{ mol KNO}_3}{2\text{ mol Pb}(\text{NO}_3)_2} \cdot \frac{1\text{ mol}}{0.0300\text{L}} = 0.29\text{ M KNO}_3$$

$$0.0200\text{L} \cdot \frac{0.20\text{ mol KI}}{1\text{L KI}} = \frac{2\text{ mol KNO}_3}{2\text{ mol KI}} \cdot \frac{1\text{ mol}}{0.0300\text{L}} = 0.26\text{ M KNO}_3$$

Therefore, KI is limiting reactant while  $\text{Pb}(\text{NO}_3)_2$  is excess reactant (LR)

Since KI is LR, 0.2 KI remains

To calculate the concentration of excess reactant remaining, we need to calculate first how many moles of  $\text{Pb}(\text{NO}_3)_2$  remain after rxn is complete. LR determines how many moles of  $\text{Pb}(\text{NO}_3)_2$  are consumed.

$$0.0200\text{L} \cdot \frac{0.20\text{ mol KI}}{1\text{L KI}} = \frac{1\text{ mol Pb}(\text{NO}_3)_2}{2\text{ mol KI}} = 0.4\text{ mol Pb}(\text{NO}_3)_2$$

How many moles of  $\text{Pb}(\text{NO}_3)_2$  did we start with?

$$0.0100\text{L} \cdot \frac{0.20\text{ mol Pb}(\text{NO}_3)_2}{1\text{L Pb}(\text{NO}_3)_2} = 0.0020\text{ mol Pb}(\text{NO}_3)_2$$

Therefore,

$$[\text{Pb}(\text{NO}_3)_2] = \frac{(0.0050\text{ mol} - 0.004\text{ mol consumed})}{0.0300\text{L (total volume)}} = 0.17\text{ M Pb}(\text{NO}_3)_2 \text{ remains}$$