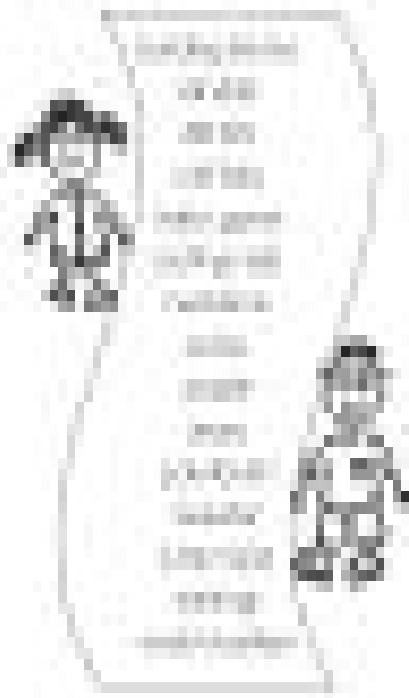


Name \_\_\_\_\_

## Practical Part

Observe the following reactions and write the balanced equation. Also write the name of the product.



Write the balanced equation for the following reactions and name the product.

1.  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$

2.  $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$

3.  $\text{Mg} + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{H}_2$

4.  $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2$

5.  $\text{Cu} + \text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{H}_2$

6.  $\text{Ni} + \text{H}_2\text{SO}_4 \rightarrow \text{NiSO}_4 + \text{H}_2$

7.  $\text{Pb} + \text{H}_2\text{SO}_4 \rightarrow \text{PbSO}_4 + \text{H}_2$

8.  $\text{Sn} + \text{H}_2\text{SO}_4 \rightarrow \text{SnSO}_4 + \text{H}_2$

9.  $\text{Co} + \text{H}_2\text{SO}_4 \rightarrow \text{CoSO}_4 + \text{H}_2$

10.  $\text{Mn} + \text{H}_2\text{SO}_4 \rightarrow \text{MnSO}_4 + \text{H}_2$

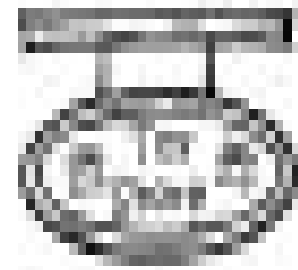


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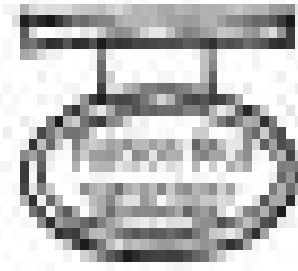


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