

Type II Binary Ionic Compounds contain Transition metals (including the Group III, IV, V, VI metals, except for Al) with non-metal ions.

Show the correct name for the following compounds.

Give correct formulas for these Type II binary compounds

copper(II) iodide	copper(I) oxide
iron(II) sulfide	manganese(IV) oxide
gold(III) chloride	gold(I) sulfide
silver bromide*	cadmium sulfide*
vanadium(V) nitride	chromium(VI) oxide
lead(IV) sulfide	zinc selenide*
gallium(I) iodide	vanadium(III) chloride
nickel(III) selenide	iron(III) bromide
silver sulfide*	tin(IV) chloride
manganese(II) fluoride	antimony(V) bromide
cobalt(III) oxide	titanium(IV) oxide

Give correct Type II names for the following compounds

FeCl ₃	Au ₂ O
FeCl ₂	AuCl ₃
V ₂ O ₅	PbO
MnO ₂	Pb ₂ S ₃
NiF ₃	Hg ₂ F ₂
CuS	HgI ₂
CuBr	CdSe
Fe ₂ O ₃	ZnBr ₂
SnF ₄	ScCl ₃
Ga ₂ S	MnCl ₂
AgI	NiCl ₂
