1. Suppose you are taking the six courses shown on the table below.

a. Complete the table.

	Units	Fraction of full load (34 units)	Letter Grade Earned	Numerical Grade Equiv.
Calculus II	6	**0.176**	B	**0.528**
Organic Chemistry 2	10	**0.295**	B	**0.885**
Microeconomics	6	**0.176**	A	**0.706**
Developmental Psychology	6	**0.176**	A	**0.706**
Piano	2	**0.063**	B	**.183**
Accounting I	6	**0.176**	C	**0.352**

________ **(sum is equal to 1)**

Total units: 34

b. Without doing any calculations, predict your quality point average (QPA) for the semester based on the grades above.

Predicted QPA: ____________________

c. Calculate your grade point average (QPA) for the semester. SHOW ALL WORK.

Add the numbers in the Numerical Grade Equiv. Column: QPA = 3.176

2. The two naturally occurring isotopes of nitrogen have masses of 14.0031 D (Dalton mass units) and 15.0001 D. The fractional abundance of these two isotopes in naturally occurring nitrogen is shown below. Based on these abundances, what would be the **atomic mass** (i.e. the average mass) of an atom of nitrogen?

Isotope	Mass	Fractional Abundance
Nitrogen-14	14.0031 D	0.993
Nitrogen-15	15.0001 D	0.007

Atomic mass = 0.993(14.0031) + 0.007(15.0001) = **14.01 D**