

**Topics:**

**Objectives:**

**Skills:**

### **Activity 1: Properties and Compounds from Chemistry**

#### **Introduction:**

Chemical reactions involve rearrangement of atoms. While the law of the conservation of mass states that the total mass of reactants must equal the total mass of products, the law of definite proportions states that the relative number of atoms of each element in a compound is always constant. This law gives us the ability to predict the properties of compounds based on their composition. The properties of a compound are determined by the elements it contains and the way the atoms interact with each other.

#### **Purposes:**

##### **Objectives:**

- (1) Know the basic properties and composition of common elements like the three gases.
- (2) Identify elements based on their atomic number or weight and the atomic weight numbers on the chart of the periodic table and their symbols.
- (3) Know the names of the gases in the periodic table.
- (4) Place the gases in a three column format.
- (5) Calculate the total mass of the three elements combining. Place your findings in a separate sheet for 24 hours.

##### **Skills:**

- (1) Calculate the mass of the gas from the total with the formula and their density and standard temperature conversion factors. However, calculate masses using the molecular and molar mass. Add this to find all the masses of the copper elements.
- (2) Add the masses of the elements. Divide the mass by the volume.
- (3) Add the masses and calculate the mass of the copper in the resulting solution.

#### **Instructions and Questions:**

- (1) Calculate the percent composition of copper in the gases.
- (2) Calculate the percent composition of zinc in the gases.
- (3) Add the masses for questions 1 & 2 and type them in the following table under Element.
- (4) Calculate the average value for the percent of copper, calculate percentages from the given lists.
- (5) Add the 10% Zinc from both. Recalling the idea of the percentage for mixing of gases for the past few years. What happens do you believe have been placed on both sides of the table that have made the discussion go on the general?