

**Type II Binary Ionic Compounds** contain Transition metals (including the Group III, IV, V, VI metals, except for Al) with non-metal ions.

Show the correct name for the following compounds.

**Give correct formulas for these Type II binary compounds**

copper(II) iodide	copper(I) oxide
iron(II) sulfide	manganese(IV) oxide
gold(III) chloride	gold(I) sulfide
silver bromide*	cadmium sulfide*
vanadium(V) nitride	chromium(VI) oxide
lead(IV) sulfide	zinc selenide*
gallium(I) iodide	vanadium(III) chloride
nickel(III) selenide	iron(III) bromide
silver sulfide*	tin(IV) chloride
manganese(II) fluoride	antimony(V) bromide
cobalt(III) oxide	titanium(IV) oxide

**Give correct Type II names for the following compounds**

FeCl <sub>3</sub>	Au <sub>2</sub> O
FeCl <sub>2</sub>	AuCl <sub>3</sub>
V <sub>2</sub> O <sub>5</sub>	PbO
MnO <sub>2</sub>	Pb <sub>2</sub> S <sub>3</sub>
NiF <sub>3</sub>	Hg <sub>2</sub> F <sub>2</sub>
CuS	HgI <sub>2</sub>
CuBr	CdSe
Fe <sub>2</sub> O <sub>3</sub>	ZnBr <sub>2</sub>
SnF <sub>4</sub>	ScCl <sub>3</sub>
Ga <sub>2</sub> S	MnCl <sub>2</sub>
AgI	NiCl <sub>2</sub>

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