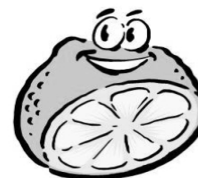


## SCIENCE 10: Chapter 5 Review Worksheet

Name: \_\_\_\_\_ Block: \_\_\_\_\_ Date: \_\_\_\_\_

1. Which of the following describes acids:
- I. Has a pH of less than 7
  - II. Can conduct electricity
  - III. Produce hydroxide ions when dissolved in water

- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III



2. What is the chemical formula for sulphurous acid?:

- A. HS
- B. HSO<sub>4</sub>
- C. H<sub>2</sub>SO<sub>3</sub>
- D. H<sub>2</sub>SO<sub>4</sub>

3. Which of the following is a base?:

- A. KCl
- B. HBr
- C. LiOH
- D. HNO<sub>3</sub>

4. What happens to the number of H<sup>+</sup> ions when HCl is added to water?:

- A. Increase
- B. Decrease
- C. Stays the same

5. What is the name for HClO<sub>3</sub>?:

- A. Chloric acid
- B. Chlorous acid
- C. Perchloric acid
- D. Hydrochlorous acid

6. What is the pH of a substance that causes methyl orange to turn yellow and methyl red to turn red?

- A. 3
- B. 6.5
- C. 4.5
- D. 8

7. How many times more acidic is a substance with a pH of 2 than a substance with a pH of 5?

- A. 10
- B. 100
- C. 1000
- D. 30

8. Complete the following table:

<i>Compound</i>	<i>Acid or Base?</i>	<i>Name</i>
H <sub>3</sub> PO <sub>4</sub>	Acid	Phosphoric acid
NH <sub>4</sub> OH	Base	Ammonium hydroxide
H <sub>2</sub> SO <sub>3</sub>	Acid	Sulfurous acid
HI	Acid	Hydroiodic acid
CH <sub>3</sub> COOH	Acid	Acetic acid
Mg(OH) <sub>2</sub>	Base	Magnesium hydroxide
HBr	Acid	Hydrobromic acid
H <sub>2</sub> CO <sub>3</sub>	Acid	Carbonic acid
NaOH	Base	Sodium hydroxide