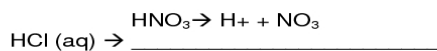


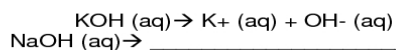
### Worksheet: Acids, Bases, and Salts Review

Name \_\_\_\_\_

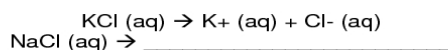
- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form \_\_\_\_\_ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form \_\_\_\_\_ ions.



- Salts are compounds that dissociate in aqueous solution releasing neither \_\_\_\_\_ ions nor \_\_\_\_\_ ions.



Using the Arrhenius definition, classify the following examples as acids, bases, or salts:

HBr _____	KCl _____
Mg(OH) <sub>2</sub> _____	H <sub>3</sub> PO <sub>4</sub> _____
HCl _____	HClO _____
KNO <sub>2</sub> _____	Al(OH) <sub>3</sub> _____
HFO <sub>4</sub> _____	KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> _____
Ba(OH) <sub>2</sub> _____	NaCl _____

Acids and bases can also be identified using an operational definition. Operational definitions are simply a list of properties.

ACIDS:

- ♦ A \_\_\_\_\_ taste is a characteristic property of all acids in aqueous solution.
- ♦ Acids react with some metals to produce \_\_\_\_\_ gas.
- ♦ Because aqueous acid solutions conduct electricity, they are identified as \_\_\_\_\_.
- ♦ Acids react with bases to produce a \_\_\_\_\_ and water.
- ♦ Acids turn \_\_\_\_\_ different colors.

BASES:

- ♦ Bases tend to taste \_\_\_\_\_ and feel \_\_\_\_\_.
- ♦ Like acids, aqueous basic solutions conduct \_\_\_\_\_, and are identified as \_\_\_\_\_.
- ♦ Bases react with \_\_\_\_\_ to produce a salt and \_\_\_\_\_.
- ♦ Bases turn \_\_\_\_\_ different colors.

Naming Acids

- Binary acids consist of \_\_\_\_\_ elements, the first being \_\_\_\_\_.
- Ternary acids consist of \_\_\_\_\_ elements. Do NOT use a prefix.  
-ate becomes \_\_\_\_\_ and -ite becomes \_\_\_\_\_

1. Give the word equation for the neutralization reaction of an acid and a base.

2. Complete these equations by creating the formula and balancing:

Hydrochloric acid + lithium hydroxide →

Acetic acid + magnesium hydroxide →

Aluminum + carbonic acid →

Magnesium + phosphoric acid →